

OFFICE OF NAVAL RESEARCH

Grant N00014-89-J-1261

R&T Code 4131038

ONR Technical Report #26

Probing the Intermolecular Hydrogen Bonding of Water Molecules at the
CCl₄/Water Interface in the Presence of Charged Soluble Surfactant
by

D. E. Gragson and G. L. Richmond

Journal of Chemical Physics, 107 (22) 1997

Department of Chemistry
1253 University of Oregon
Eugene, OR 97403

June 1998

Reproduction in whole, or in part, is permitted for any purpose of the United States Government.

This document has been approved for public release and sale; its distribution is unlimited.

19980623 045

REPORT DOCUMENTATION PAGE

Form Approved
OMB No. 0704-0188

1. AGENCY USE ONLY (Leave Blank)		2. REPORT DATE 30 May 1998		3. REPORT TYPE AND DATES COVERED Technical 6/1/97-5/31/98	
4. TITLE AND SUBTITLE Probing the Intermolecular Hydrogen Bonding of Water Molecules at the CCl ₄ /Water Interface in the Presence of Charged Soluble Surfactant				5. FUNDING NUMBERS N00014-89-J-1261	
6. AUTHOR(S) D. E. Gragson and G. L. Richmond					
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) Dept. of Chemistry University of Oregon Eugene, OR 97403				8. PERFORMING ORGANIZATION REPORT NUMBER ONR Technical Report #26	
9. SPONSORING/MONITORING AGENCY NAME(S) AND ADDRESS(ES) Dr. Peter Schmidt Office of Naval Research Physical Science and Technology, ONR 331 800 North Quincy Street Arlington, VA 22217-5000				10. SPONSORING/MONITORING AGENCY	
11. SUPPLEMENTARY NOTES Journal of Chemical Physics, 107 (22) 1997					
12A. DISTRIBUTION / AVAILABILITY STATEMENT Approved for public release: distribution unlimited				12B. DISTRIBUTION CODE	
13. ABSTRACT (Maximum 200 words) Please see attached abstract					
14. SUBJECT TERMS Molecular Structure and Hydrogen Bonding of Water Molecules at the CCl ₄ /Water				15. NUMBER OF PAGES 14	
				16. PRICE CODE	
17. SECURITY CLASSIFICATION OF REPORT Unclassified	18. SECURITY CLASSIFICATION OF THIS PAGE Unclassified	19. SECURITY CLASSIFICATION OF ABSTRACT Unclassified	20. LIMITATION OF ABSTRACT		

Probing the Intermolecular Hydrogen Bonding of Water Molecules at the CCl₄/Water Interface in the Presence of Charged Soluble Surfactant

D. E. Gragson and G. L. Richmond

University of Oregon, Department of Chemistry, Eugene, OR 97403

Abstract

The molecular structure and hydrogen bonding of water molecules at the CCl₄/water interface in the presence of a charged soluble surfactant has been explored in this study using vibrational sum frequency generation. By examining OH stretching modes that are highly sensitive to the local hydrogen bonding environment it is found that water molecules at the CCl₄/water interface both in the presence and absence of a charged soluble surfactant are predominantly in a tetrahedral arrangement much like the structure of ice. Isotopic dilution studies have been employed to further characterize this ice-like interfacial structure. A blue-shift of approximately 120 cm⁻¹ of the ice-like OH stretching mode is observed upon dilution with D₂O. The first vibrational spectra of the OH stretching mode from uncoupled HOD molecules at the CCl₄/water interface is also reported.

The interactions between water and charged amphiphiles at surfaces can play a dominant role in many processes including micellization, stabilization of membranes and protein activity. At an oil/water interface an amphiphile alters the hydrogen bonding of water at the junction, resulting in a lowering of the interfacial tension. How these changes in hydrogen bonding with surfactant adsorption are manifested in the nonlinear vibrational spectroscopy of interfacial water at the $\text{CCl}_4/\text{H}_2\text{O}$ interface is the focus of this paper.

Total internal reflection sum frequency generation (TIR-VSFG) is employed here in a unique set of experiments which characterize in detail the hydrogen bonding between interfacial water molecules at the liquid/liquid interface in the presence of a simple charged soluble surfactant. The combination of the surface sensitivity of VSFG and the enhancement afforded by operating in a total internal reflection geometry makes TIR-VSFG a powerful tool in studies such as these. The spectroscopy of the OH stretching modes of interfacial water molecules has been used in this study as a diagnostic of interfacial water structure and hydrogen bonding. These modes are particularly sensitive to the hydrogen bonding environment of water molecules as shown by numerous IR and Raman studies of bulk water.^{6,7} For bulk water, these studies have shown that water molecules partition into ice-like and water-like structures. We report that at the CCl_4 /water interface both in the presence and in the absence of surfactant, nearly all interfacial molecules are in a tetrahedral arrangement much like the structure of ice with little or no evidence for a water-like structure.

Further explorations of the ice-like character of the interfacial water molecules in the presence of surfactants is obtained through isotopic dilution experiments. Isotopic dilution studies in which the OH stretching vibrations are progressively decoupled by additions of D_2O have been invaluable in studies of hydrogen bonding in bulk water⁸ and ice⁹ but this is the first study in which such experiments have been applied to liquid surfaces. The shift in energy observed in the ice-like O-H peak with uncoupling of the OH symmetric stretching vibration has provided us with a quantitative means of comparing the hydrogen bonding at this interface with that measured for bulk ice and super cooled liquid water. Through isotopic dilution studies we have also obtained the first vibrational spectra of the uncoupled OH stretching mode from HOD molecules at the

CCl_4 /water interface. This mode has been shown to be an excellent indicator of the local hydrogen bonding environment in bulk studies.¹⁰

Figure 1 shows the VSFG spectra from the CCl_4 /water interface with the aqueous phase having a bulk sodium dodecyl sulfate (SDS) concentration of 1.0 mM. All spectra presented were obtained under S_{sfg} , S_{vis} , P_{ir} polarization conditions which picks out vibrational modes with components of the transition dipole moment perpendicular to the plane of the interface. The series of spectra were taken with various mixtures of H_2O and D_2O ranging from a mole fraction of 1.00 to 0.10. The mole fraction of H_2O , D_2O , and HOD were calculated by assuming complete isotopic exchange.¹⁰ The presence of SDS at the interface is evidenced by the C-H stretching modes of the alkyl chain and is also verified by interfacial tension measurements. Because these modes have been examined in detail in previous studies^{3,11} from this laboratory under higher spectral resolution, further discussion is not warranted here. The shoulder in the spectra near 3000 cm^{-1} has not been studied in any detail here but could be due to the OH stretch from water molecules that are hydrogen bonded to the charged sulfate head group with the large red-shift presumably a result of the strong hydrogen bond. The solid lines in Figure 1 are a spectral fit to the data from which we are able to extract peak intensities, positions, bandwidths and integrated areas.

Focusing on the spectrum in Figure 1 corresponding to SDS in 100% H_2O , the dominant feature in the OH stretching spectral region is centered around 3200 cm^{-1} . We attribute the observed peak at 3200 cm^{-1} to the OH-SS-S mode which corresponds to the in-phase coupled symmetric stretch from water molecules that have complete tetrahedral coordination much like the structure of ice. In contrast, for SDS at the air/water interface the VSFG spectrum shows two peaks,¹¹ one corresponding to the OH-SS-S mode and an additional peak located between 3400 cm^{-1} to 3450 cm^{-1} which is attributed to water molecules that have incomplete tetrahedral coordination and a more water-like structure. These two peaks are commonly referred to as the "ice-like" and "water-like" modes with the ice-like mode corresponding to a higher degree of hydrogen bond order and the water-like mode to a lower degree of hydrogen bond order.^{5,8} From

the dominance of the ice-like mode in the VSFG spectra from SDS at the $\text{CCl}_4/\text{H}_2\text{O}$ interface we infer that the prevailing structure of the interfacial water molecules in the presence of a charged surfactant is a tetrahedral arrangement much like the structure of ice. As will be described in a later publication¹², relative to the neat $\text{CCl}_4/\text{H}_2\text{O}$ interface, addition of surfactant results in increased intensity in the ice-like mode due to electrostatic field induced orientation of water molecules in the double layer region.

Another OH mode often observed in bulk and surface water studies⁵ but not observed here is located at approximately 3680 cm^{-1} and is attributed to the free OH stretch from water molecules that have hydrogen atoms not participating in hydrogen bonding. The laser system employed limits our ability to detect this mode.¹³ The red-shift of the peak position of the OH stretching modes with increased intermolecular hydrogen bonding as is seen in the three aforementioned modes has been thoroughly examined.^{7,14} The shift occurs because hydrogen bonding actually “steals” bond strength from the OH bond and thus stronger hydrogen bonds result in weaker OH covalent bonds and vibrational modes at lower energy. A comparison of the peak positions with the degree of hydrogen bonding illustrates the well known trend that the peak position of the OH stretching mode is red-shifted with increasing hydrogen bonding. Accompanying the red-shift of the peak frequency with increased hydrogen bonding is a large increase in the bandwidth of the OH stretch. This increase in the bandwidth results from dynamic dipole-dipole coupling between neighboring water molecules which produces a distribution of low and high frequency stretching modes.^{7,14,15} The nature of this distribution also has an effect on the position of the peak frequency of the OH stretching mode. Deconvolution of these two effects, hydrogen bonding and intermolecular coupling, on the energetics of the OH stretching peaks in the vibrational spectra is difficult and generally requires the study of HOD in H_2O or D_2O which eliminates the intermolecular coupling effect. However, the extent of hydrogen bonding can be inferred through a comparison of the relative number of water molecules contributing to the ice-like and water-like

modes. This comparison is possible since the ice-like mode is indicative of more complete hydrogen bonding than the water-like mode.

Our observation that at the CCl_4 /water interface the ice-like mode dominates the OH stretching spectral region provides direct evidence that there is more complete hydrogen bonding between neighboring water molecules at the interface than in the bulk. This observation is a consequence of the hydrophobic effect which causes the interfacial water molecules to rearrange into a tetrahedral structure thus maximizing the intermolecular hydrogen bonding in the presence of a nonpolar solute such as CCl_4 .^{1,12,16} We are probing specifically the asymmetric interfacial region which for an ionic strength of 1.0 mM corresponds to at most 30 water layers [Gragson, 1997 #14], thus one might expect CCl_4 to be present in this region. In the bulk aqueous phase the water molecules are not influenced by the presence of a non-polar molecule and the water-like peak accompanies the ice-like peak. At the *neat* CCl_4 /water interface we also find that the ice-like mode dominates the VSFG spectrum.¹⁷ The similarity of the water structure at the CCl_4 /water interface both in the presence and in the absence of a charged soluble surfactant further allows us to infer that the presence of the surfactant and counter ions in the aqueous phase does not disrupt to any measurable degree the hydrogen bond ordering of the interfacial water molecules.

Further information about the interfacial water hydrogen bonding structure can be extracted from the spectra of the isotopic mixtures shown in Figure 1 and Figure 2 where the latter corresponds to spectra from the lowest H_2O concentrations depicted on an expanded scale for clarity. Numerous IR and Raman studies of $\text{H}_2\text{O}/\text{D}_2\text{O}$ mixtures have been conducted in the past to gain a better understanding of the structure of water in both bulk liquid water and bulk ice.^{6,7,9} As D_2O (H_2O) is added to H_2O (D_2O) the intermolecular coupling between the OH (OD) oscillators decreases as a result of the difference in hydrogen bonding between H_2O and D_2O and the difference in energy of the OH and OD stretches. The intermolecular decoupling as D_2O (H_2O) is added to H_2O (D_2O) manifests itself in the IR and Raman spectra as a blue-shift in the spectral position of the OH (OD) stretching vibrations. This blue-shift occurs as the mole fraction of H_2O (D_2O) is decreased and results from the decreased intermolecular coupling of the OH oscillators.

Experiments conducted on cubic and amorphous ice⁹ as well as super cooled liquid water¹⁵ have shown a blue-shift in the OH-SS-S peak of approximately 100-120 cm^{-1} as the mole fraction of H_2O varies from 1.00 to 0.01.

Close inspection of the spectra in Figure 1 shows a similar blue-shift of the OH-SS-S peak position with decreasing H_2O mole fraction. The peak position goes from 3200 cm^{-1} at a mole fraction of 1.00 to 3320 cm^{-1} at a mole fraction of 0.05. The blue shift in the peak frequency of the OH-SS-S with decreasing H_2O mole fraction is a result of the intermolecular decoupling of the OH oscillators by the addition of OD oscillators as previously mentioned. The magnitude of the blue shift that we observe for the $\text{CCl}_4/\text{H}_2\text{O}$ -SDS system is approximately the same, 120 cm^{-1} , as has been observed for both bulk amorphous ice and cubic ice⁹ as well as super cooled water at -5°C .¹⁵ This similarity provides further support that the interfacial water molecules are in an ice-like arrangement with a high degree of hydrogen bond order and a tetrahedral structure. Figure 3 shows the peak frequency of the OH-SS-S stretch obtained from the fits to the spectral data plotted as a function of the mole fraction of H_2O . From Figure 3 we see that the peak frequency of the OH-SS-S mode blue-shifts in nominally a linear fashion with decreasing H_2O mole fraction over the range from 1.0 to 0.2. Beyond a mole fraction of 0.2 the slope of the frequency shift with decreased H_2O becomes significantly steeper. This observation is most likely due to the increased HOD component present at the interface interfering with the fitting of the spectra to the uncoupled OH stretch from H_2O .

Figure 2 shows the VSFG spectra from the lowest mole fraction H_2O solutions studied. At a H_2O mole fraction of 0.108 a small shoulder is observed at approximately 3460 cm^{-1} on the high frequency side of the blue-shifted OH-SS-S peak. We attribute this peak to the uncoupled OH stretch (OH-S) from interfacial HOD molecules produced by isotopic exchange between H_2O and D_2O . The VSFG spectrum from the 0.05 mole fraction H_2O solution shows that this peak actually dominates the OH stretching spectral region due to the very small concentration of H_2O and the much larger (0.35 mole fraction) HOD concentration. The absolute intensity from the OH stretch of HOD is much weaker than the OH stretch of H_2O due to diminished hydrogen bonding and

uncoupling of the OH stretching vibration of interfacial HOD molecules. HOD is both intermolecularly and intramolecularly uncoupled and thus has been extensively used in the determination of structural characteristics from vibrational spectra.^{7,10,18} This work represents the first observation of the OH stretch from uncoupled HOD molecules at a liquid/liquid interface in the presence of a charged soluble surfactant.

Studies using polarized Raman spectroscopy have suggested a two species model for the hydrogen bonded HOD molecules in bulk water solutions.⁸ While more complicated models are currently employed, the two species model is adequate for the analysis here. The first species is composed of HOD molecules with two equivalent strong hydrogen bonds while the second is composed of HOD molecules with one weak and one strong hydrogen bond. The OH stretch from the more strongly hydrogen bonded HOD in D₂O has been shown to occur between 3300 cm⁻¹ and 3500 cm⁻¹ and is very temperature sensitive, shifting to higher frequencies with increased temperature. The peak position of the more strongly hydrogen bonded OH(OD) stretching mode in the liquid state varies from 3435 cm⁻¹ to 3489 cm⁻¹ (2520 cm⁻¹ to 2568 cm⁻¹) as the temperature varies from 10° to 90° C.⁸ In contrast, the peak position of the more weakly hydrogen bonded OH(OD) from HOD in D₂O(H₂O) at approximately 3600 cm⁻¹ (2650 cm⁻¹) shows minimal temperature dependence.⁸ That the OH stretch from interfacial HOD molecules at room temperature is located at approximately 3460 cm⁻¹ agrees well with the position of the more strongly hydrogen bonded OH stretch. Further experiments are in progress to explore the temperature dependence of this mode.

Our VSFG studies of the hydrogen bonding and structure of water molecules at the CCl₄/water interface in the presence of a charged soluble surfactant provide new insight into differences between interfacial and bulk water molecules. We find that water molecules at the CCl₄/water interface both in the presence and absence of a charged soluble surfactant are predominantly in a tetrahedral arrangement much like the structure of ice whereas water molecules in the bulk are equally partitioned into water-like and ice-like modes at room temperature. Isotopic dilution studies have been used for the first time to understand the hydrogen bonding of water

molecules at a liquid surface. A blue-shift of approximately 120 cm^{-1} in the ice-like peak is observed upon dilution which we attribute to intermolecular uncoupling of the OH oscillators. The similarity in the magnitude of this frequency shift with that of bulk amorphous and cubic ice as well as super cooled liquid water further confirms our conclusion that the interfacial water molecules are in an ice-like arrangement. At low H_2O mole fractions we find that we are sensitive to the OH stretching mode from uncoupled HOD at the $\text{CCl}_4/\text{water}$ interface. This mode has been shown to be essential in the full characterization of the structure of bulk water in both liquid and solid phases from IR and Raman spectroscopic studies. Our future endeavors will include exploiting this mode to probe the interfacial hydrogen bonding environment as a function of temperature and the presence of structure making and breaking ions as well as to probe structural dependencies on the interfacial potential.

Acknowledgments

The authors would like to thankfully acknowledge the Office of Naval research and the National Science Foundation (CHE-9416856) for support of this work.

References

- ¹Malcom N. Jones and Dennis Chapman, *Micelles, Monolayers, and Biomembranes* (Wiley-Liss, New York, 1995).
- ²G. R. Bell, C. D. Bain, and R. N. Ward, *J. Chem. Soc., Faraday Trans.* **92** (4), 515-523 (1996).
- ³J. C. Conboy, M. C. Messmer, and G. L. Richmond, *J. Phys. Chem.* **100** (18), 7617-7622 (1996).
- ⁴D. E. Gragson, B. M. McCarty, and G. L. Richmond, *J. Phys. Chem.* **100** (34), 14272-14275 (1996).
- ⁵Quan Du, Eric Freysz, and Y. Ron Shen, *Science* (Washington, D. C.) **264** (5160), 826-8 (1994).
- ⁶G. E. Walrafen, "Water: A Comprehensive Treatise," in *Water: A Comprehensive Treatise*, edited by F. Franks (Plenum Press, New York, 1972), Vol. 1, pp. 151-254.
- ⁷D. Eisenberg and W. Kauzmann, *The Structure and Properties of Water* (Oxford University Press, New York, 1969).
- ⁸J. R. Scherer, M. K. Go, and Saima Kint, *J. Phys. Chem.* **78** (13), 1304-13 (1974).
- ⁹J. P. Devlin, *J. Chem. Phys.* **90** (3), 1322-29 (1989).
- ¹⁰J. Wiafe-Akenten and R. Bansil, *J. Chem. Phys.* **78** (12), 7132-37 (1983).
- ¹¹D. E. Gragson, B. M. McCarty, and G. L. Richmond, *J. Am. Chem. Soc.* **119** (26), 6144-52 (1997).
- ¹²D. E. Gragson and G. L. Richmond, *J. Am. Chem. Soc.*, Submitted (1997).
- ¹³D. E. Gragson, B. M. McCarty, G. L. Richmond *et al.*, *J. Opt. Soc. Am. B* **13** (9), 2075-2083 (1996).
- ¹⁴E. Whalley, *Can. J. Chem.* **55**, 3429-41 (1977).
- ¹⁵J. L. Green, A. R. Lacey, and M. G. Sceats, *Chem. Phys. Lett.* **130** (1,2), 67-71 (1986).

¹⁶Charles Tanford, *The Hydrophobic Effect* (Wiley-Interscience, New York, 1973).

¹⁷D. E. Gragson and G. L. Richmond, *Langmuir* **In Press** (1997).

¹⁸T. T. Wall and D. F. Horning, *J. Chem. Phys.* **43** (6), 2079-87 (1965).

Figure Captions

Figure 1

VSFG spectra from the $\text{CCl}_4/\text{H}_2\text{O}$ interface with 1.0 mM SDS in the aqueous phase as a function of mole fraction of H_2O under S-sf, S-vis, P-ir polarization conditions. Solid lines are a spectral fit to the data.

Figure 2

VSFG spectra from the $\text{CCl}_4/\text{H}_2\text{O}$ interface with 1.0 mM SDS in the aqueous phase for low mole fractions of H_2O under S-sf, S-vis, P-ir polarization conditions.

Figure 3

Plot of the peak position of the OH-SS-S mode from the spectral data fits as a function of the mole fraction of H_2O .

Figure 1

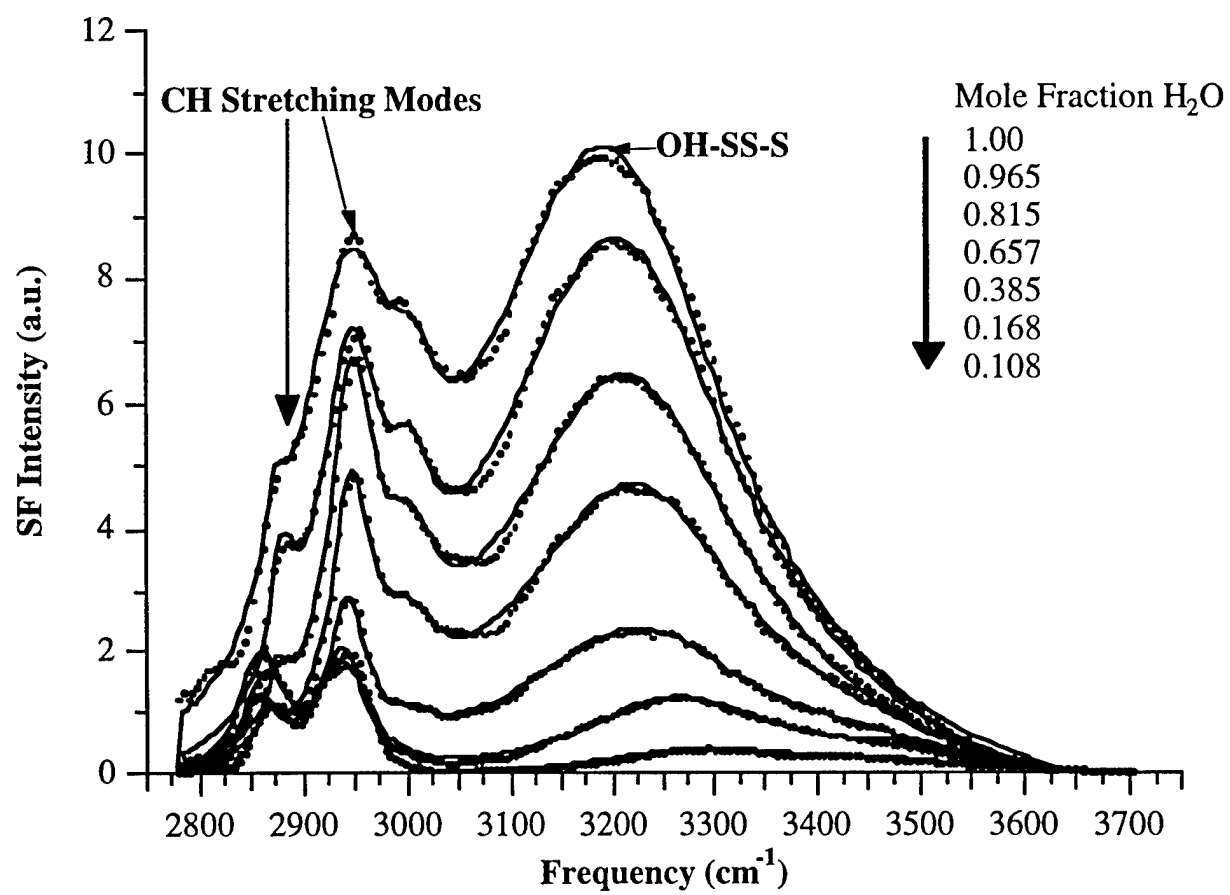


Figure 2

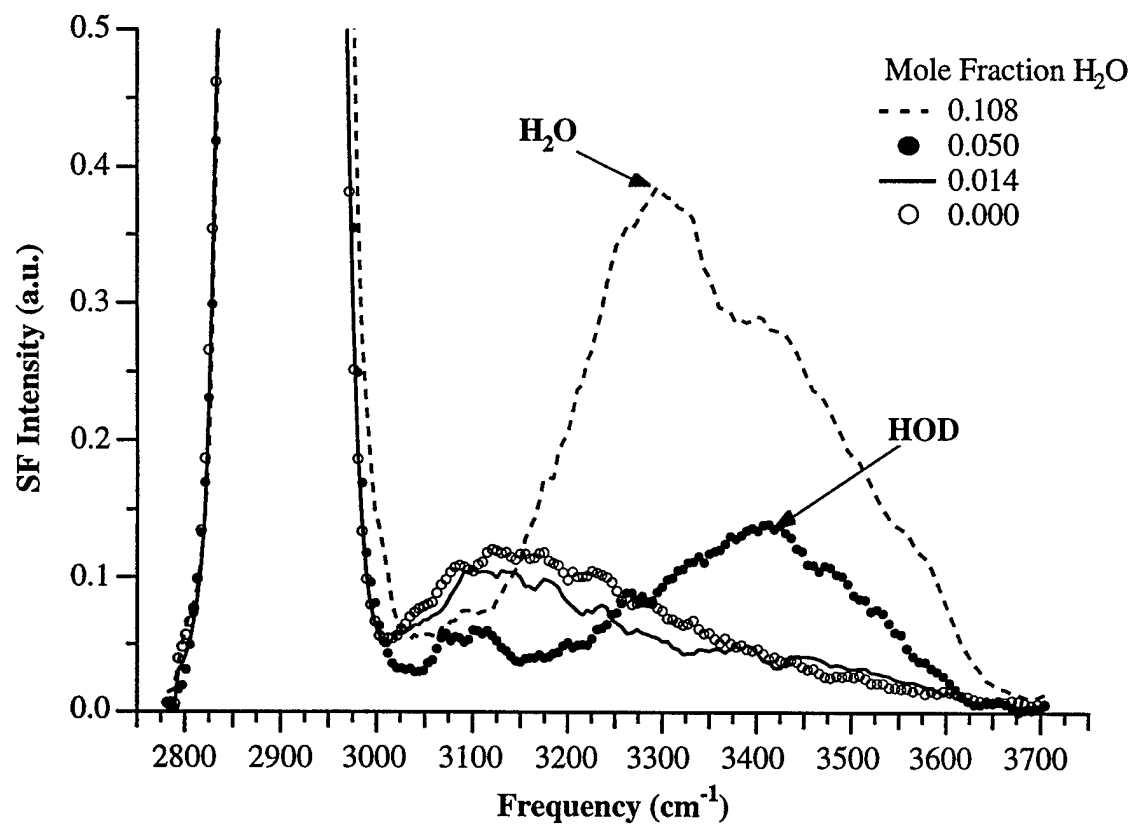


Figure 3

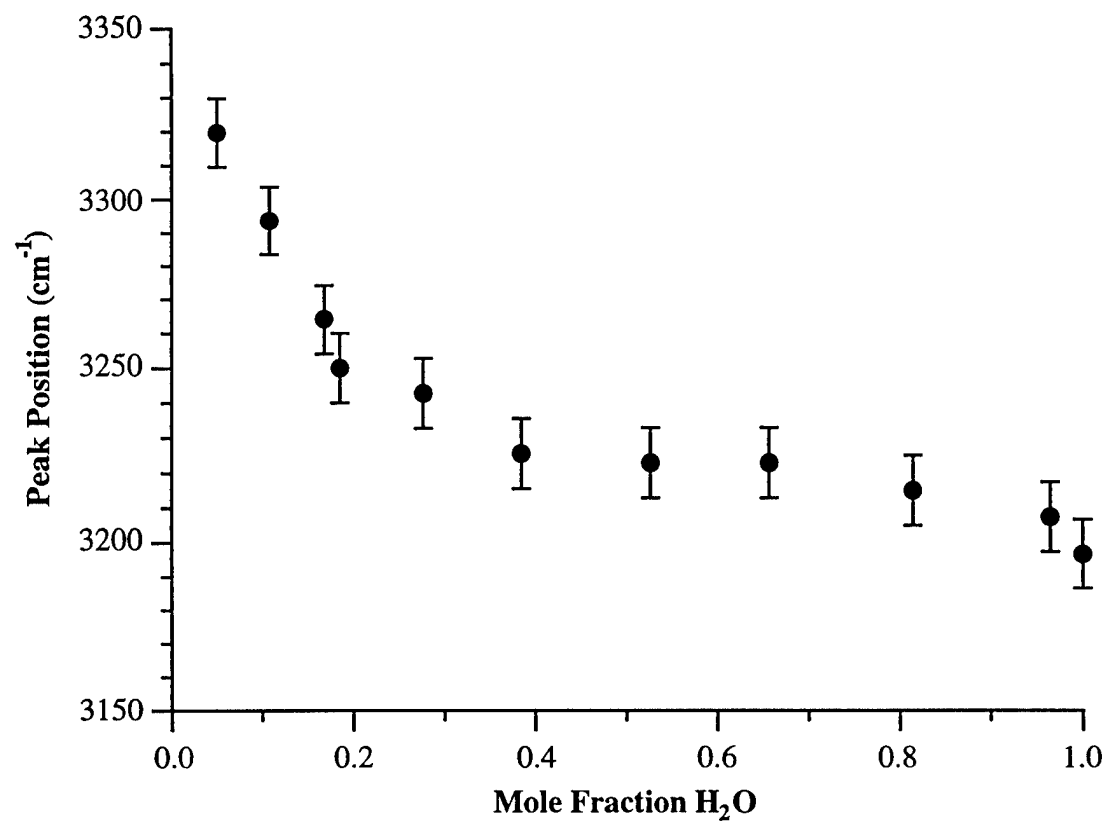


Figure 2

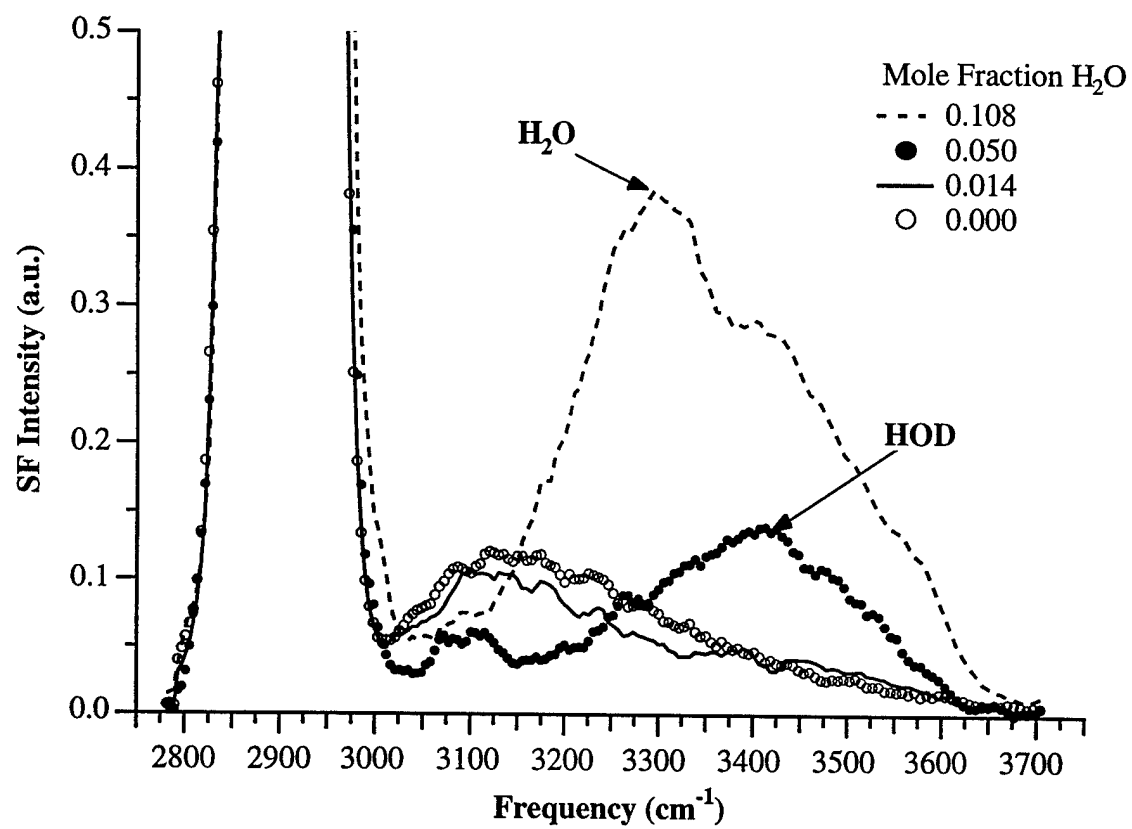


Figure 3

